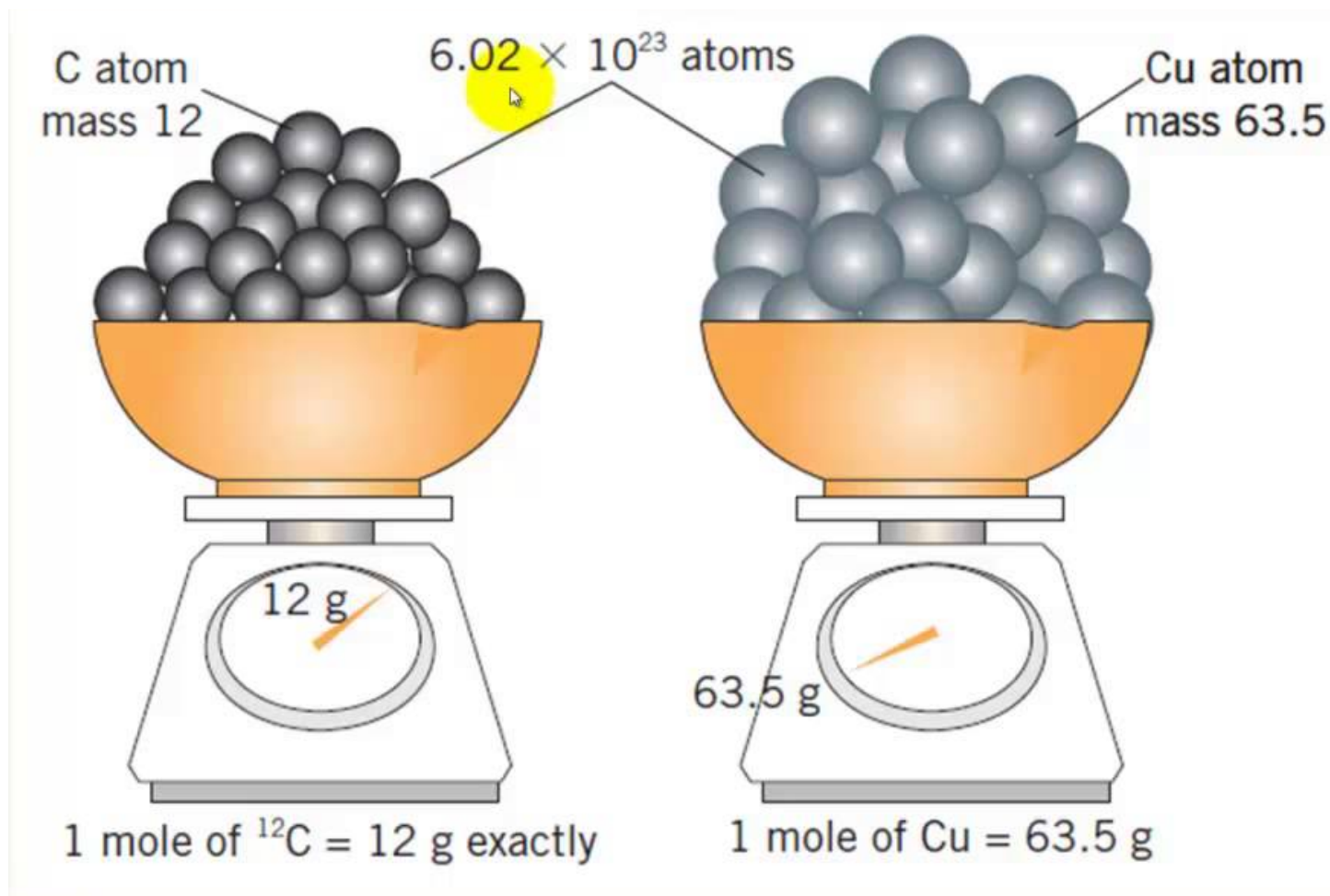


El **mol** d'una entitat (àtom, molècula, ió, electró, o qualsevol altra partícula o grup de partícules) és la quantitat de substància continguda en **$6,02214 \cdot 10^{23}$** entitats.



$$M_r[\text{C}] = 12$$
$$M_c = 12 \text{ g/mol}$$

$$M_r[\text{Cu}] = 63.5$$
$$M_c = 63.5 \text{ g/mol}$$

$$M_r[C]=12$$
$$M_C=12 \text{ g/mol}$$

$$M_r[Na]=23$$
$$M_C=23 \text{ g/mol}$$

$$M_r[Cl]=35,5 \quad M_r[Cl_2]=71$$
$$M_{Cl_2}=71 \text{ g/mol}$$

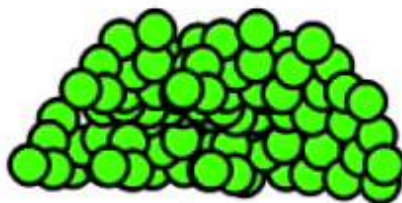
1.00mol C



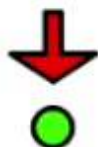
is
12.011g
and
contains
 6.02×10^{23}
of these



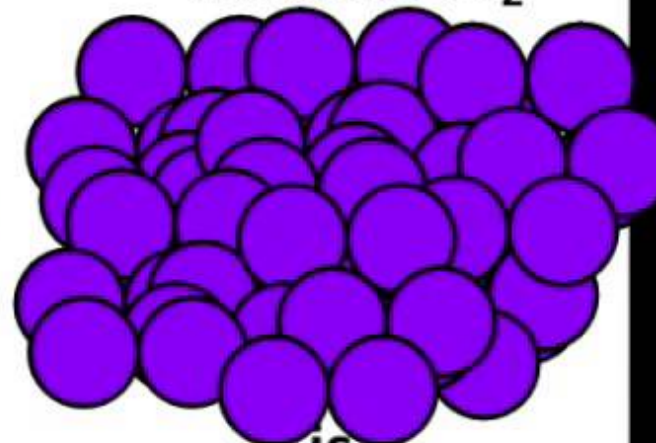
1.00mol Na



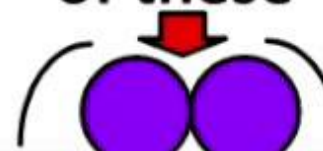
is
22.990g
and
contains
 6.02×10^{23}
of these



1.00mol Cl₂



is
70.906g
and
contains
 6.02×10^{23}
of these



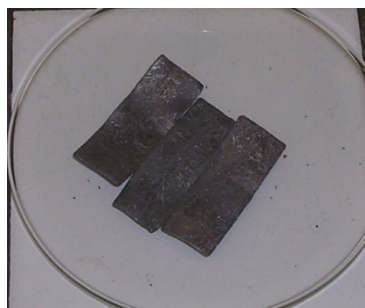
El **mol** d'una entitat (àtom, molècula, ió, electró, o qualsevol altra partícula o grup de partícules) és la quantitat de substància continguda en **$6,02214 \cdot 10^{23}$** entitats.

$$M_r[S]=32$$
$$M_S=32 \text{ g/mol}$$



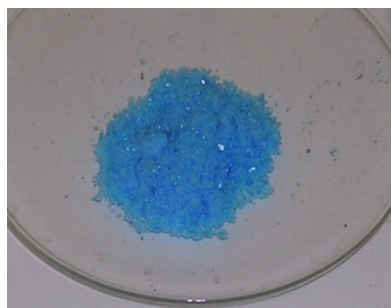
0,1 mol de S
3,2 g

$$M_r[Pb]=207$$
$$M_{Pb}=207 \text{ g/mol}$$



0,1 mol de Pb
20,7 g

$$M_r[CuSO_4]=250$$
$$M_{CuSO_4}=250 \text{ g/mol}$$



0,1 mol de $CuSO_4$
25 g

$$M_r[Cu]=64$$
$$M_{Cu}=64 \text{ g/mol}$$



0,1 mol de Cu
6,4 g

$$M_r[H_2O]=18$$
$$M_{H_2O}=18 \text{ g/mol}$$



0,1 mol de H_2O
1,8 g

1 mol conté $6,022 \cdot 10^{23}$ partícules

$$N_A = 6,022 \cdot 10^{23} \text{ mol}^{-1}$$